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## THE IMPURITY ION INFLUENCE ON THE OPTICAL AND PHOTOCATALYTIC PROPERTIES OF ANATASE AND RUTILE

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*The influence of Cu, Fe, Co, and Cr transition metal ion impurities on the optical absorption spectra of polydisperse powders of nanocrystalline TiO<sub>2</sub> rutile and anatase synthesized in identical experimental conditions is studied. For all rutile samples, no spectral shift of the absorption edge is observed, while a significant red shift of the absorption edge for modified anatase is detected. The shift value correlates well with the relative position of the impurity levels in the TiO<sub>2</sub> band gap. The effect of the impurities on the rutile optical absorption is observed only in the region 3,0-4,5 eV, whereas the modified anatase absorption changes significantly in the region 3,0-5,5 eV. The band gap remains practically unchanged for modified rutile, whereas it decreases for modified anatase, being the largest for A/Cr and A/Fe samples. The photocatalytic activity of rutile and anatase samples in the reaction of photodestruction of organic dye safranin under UV irradiation is investigated. It is shown that the transition metal impurities in rutile, at their high concentration, become additional recombination centers for photoexcited electron-hole pairs that leads to the inhibition of photocatalysis. Increase in the rate of photocatalysis in the presence of modified anatase correlates well with the increasing optical absorption of modified anatase in the region 4,0-5,5 eV.*

**Keywords:** TITANIUM DIOXIDE, ANATASE, RUTILE, TRANSITION METAL IMPURITIES, OPTICAL PROPERTIES, PHOTOCATALYSIS.

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### 1. INTRODUCTION

Titanium dioxide exhibits interesting optical properties, such as high transparency in the visible spectrum range, high index of refraction (2,4-2,7), band gap width of TiO<sub>2</sub> is equal to 3,0-3,2 eV depending on the crystalline structure. These properties allow to efficiently use TiO<sub>2</sub> in photocatalysis, solar cells, sensors, and for photon crystals. But wide band gap of TiO<sub>2</sub> allows to activate it only in the ultraviolet (UV) light that restricts the possible applications.

Many investigations have been carried out concerning the influence of different impurity atoms and compounds on the electron and optical properties of  $\text{TiO}_2$  for the purpose of the improvement of these properties and extension of  $\text{TiO}_2$  operating range to the visible spectrum region [1-3]. Anpo et al. [1] have shown that implantation of transition metal ions into  $\text{TiO}_2$  shifts its optical absorption edge toward lesser energies. In spite of a large amount of publications devoted to the effect of impurities on the optical and electron properties of  $\text{TiO}_2$ , it is quite difficult to make a direct comparison between the experimental results of different authors because of different conditions of  $\text{TiO}_2$  synthesis and experimental techniques. Synthesis conditions, thermal and chemical treatment of  $\text{TiO}_2$ , type and concentration of impurities are the determinative parameters in the given case.

The aim of this work was to investigate the influence of Cu, Fe, Co, Cr transition metal impurities on the optical and photocatalytic properties of polydisperse nanocrystalline  $\text{TiO}_2$  of rutile (R) and anatase (A) modification with a high degree of chemical purity synthesized in the same experimental conditions. Optical absorption of  $\text{TiO}_2$  samples were studied in a wide UV and visible spectrum range from 220 to 730 nm (in the photon energy region of 1,75-5,5 eV). Photocatalytic activity of the samples was investigated in the reaction of photodestruction of organic dye safranin under UV irradiation in the region of 254 nm. The obtained experimental results are presented and discussed in the given work.

## 2. EXPERIMENT

### 2.1 Materials

Polydisperse pure titanium dioxide with nanocrystalline structure of rutile or anatase was synthesized by thermal hydrolysis of  $\text{TiCl}_4$  chloride solutions in the presence of colloidal nucleuses of anatase or rutile [4]. This method allows to obtain high purity  $\text{TiO}_2$  with impurity (Cu, Fe, Co, Cr, Mn, V, Ni) content of  $10^{-5}$ - $10^{-6}$  mas%. Then obtained samples of nanocrystalline anatase or rutile were modified by separate transition metal ions  $\text{Cu}^{2+}$ ,  $\text{Fe}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Cr}^{3+}$  by adsorption from dilute aqueous solutions of the corresponding salts at pH = 7-8. Adsorption of impurity cations was performed during 24 hours to achieve the sorption equilibrium. Then samples were filtered out, dried at the temperature of 120°C, thoroughly washed by bidistilled water (until total absence of impurity ions in filtrate), and dried in the air at the temperature of 300°C. Samples for optical investigations were produced as compressed to a tablet  $\text{TiO}_2$  and KBr powder mixtures ( $\text{TiO}_2$  content was equal to 0,1%).

### 2.2 Investigation methods

Structural properties and phase composition of  $\text{TiO}_2$  samples were studied by the X-ray diffractometry (XRD) method on the diffractometer DRON-2 with  $\text{Cu}_{K\alpha}$  radiation ( $\lambda = 0,154$  nm). Chemical composition of the samples was investigated by the method of the X-ray fluorescence analysis (XRF) using the device XNAT-Control. The scanning electron (SEM) microscope JEOL JSM 6490 with X-ray spectrometer was used to explore the surface morphology and for elemental analysis of the samples. SEM images and energy-dispersion spectra (EDS) were obtained at the voltage of 20 kV.

Optical absorption spectra of TiO<sub>2</sub> samples in the UV and visible spectrum range (in the wavelength range  $\lambda = 220\text{-}750$  nm) were studied using the multi-channel optical analyzer SOLAR SL40-2 (3648-pixel CCD-detector TCD1304AP, diffraction grating is 600 mm<sup>-1</sup>, spectral resolution is 0,3 nm, operating speed is 7 ms) under control of portable PC. Deuterium lamp DDC-30 ( $\lambda_{1\text{max}} = 245$  nm and  $\lambda_{2\text{max}} = 311$  nm) and xenon lamp ( $\lambda_{\text{max}} = 472$  nm) were used as the sources of UV and visible light, respectively. Measurements were performed in standard conditions at room temperature.

Photocatalytic activity of the samples was investigated using the model reaction of photocatalytic destruction of organic dye safranin (C<sub>20</sub>H<sub>19</sub>ClN<sub>4</sub>: H<sub>2</sub>O solution) in a quartz reactor under the action of UV irradiation of the mercury lamp BUV-30 ( $\lambda_{\text{max}} = 254$  nm, power 30 W). Changes of the safranin concentration in solution in certain time intervals were determined using absorption spectrophotometer SPECORD 20 UV-VIS.

### 3. RESULTS AND DISCUSSION

#### 3.1 Structural and elemental analysis of TiO<sub>2</sub> samples

Structural characteristics of the investigated TiO<sub>2</sub> samples are presented in Table 1. In conformity with the XRD analysis results, average nanocrystal sizes in polydisperse TiO<sub>2</sub> powders are equal to 10-36 nm for rutile and 8-16 nm for anatase.

*Table 1 – Characteristics of the synthesized TiO<sub>2</sub> samples*

TiO <sub>2</sub> samples	Impurity content, at% (XRF)*	Impurity content, at% (XRF)**	Impurity content, at% (EDS)**	Crystal grain size, nm (XRD)	Cell parameter <i>a</i> , Å (XRD)	Cell parameter <i>c</i> , Å (XRD)
<b>RUTILE</b>						
R	–	–	–	21	4,591	2,968
R/Cr	2,8	1,0	2,3	36	4,591	2,968
R/Co	7,0	2,4	1,1	29	4,564	2,981
R/Fe	6,3	2,2	3,5	15	4,550	3,000
R/Cu	12,5	4,5	4,6	10	4,553	3,026
<b>ANATASE</b>						
A	–	–	–	16	3,775	9,402
A/Cr	7,5	2,6	–	15	3,775	8,983
A/Co	9,3	3,3	–	13	3,781	9,099
A/Fe	7,5	2,6	–	9	3,787	9,015
A/Cu	15,3	5,7	–	8	3,793	9,343

\* per one Ti atom;

\*\* per one atom of TiO<sub>2</sub> compound.

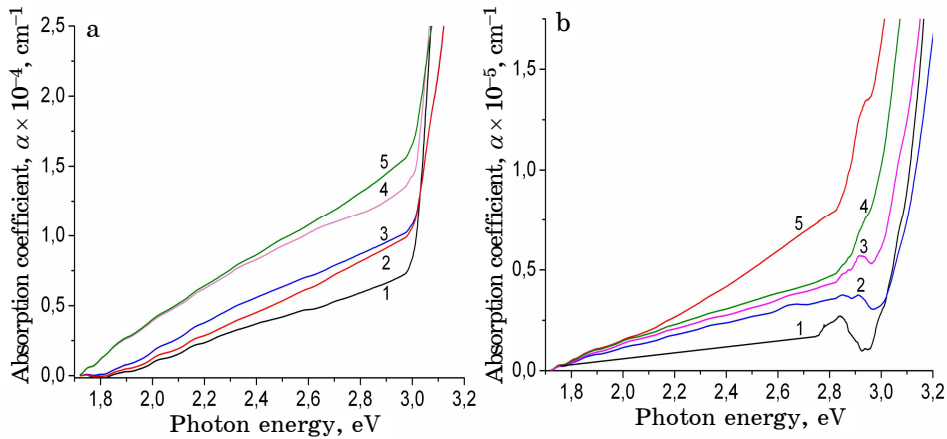
Impurity ion content in synthesized samples was estimated using the XRF and EDS analysis methods. Elemental composition of the samples (content of Ti atoms, impurity ions, oxygen as well) was determined by the EDS data. Results of EDS for modified rutile were obtained using microprobe analysis realized in different surface points of TiO<sub>2</sub> particles. We have to note that by the XRF data, impurity concentrations are calculated per total amount of

Ti atoms, while EDS is based on total amount of Ti atoms, impurity atoms, and oxygen. Results of the XRF analysis were re-calculated per total atomic content in  $\text{TiO}_2$  for the comparison with the EDS data.

Based on the XRF and EDS data, an arrangement of impurity cations in rutile samples is analyzed. Thus, Cu atoms are uniformly distributed both in the particle volume and on the surface; Cr and Fe atoms are more segregated on the surface of  $\text{TiO}_2$  particles, and Co atoms are more incorporated to the particle volume. By the EDS data, content of Cr and Fe atoms in rutile is 1,5-2 times more than by the XRF data. Such distribution of Cr and Fe atoms on the surface and in the volume of  $\text{TiO}_2$  particles is explained by their low solubility (the limit is about 1 mas% ) in  $\text{TiO}_2$  [5]. We have to note, content of Co ions on the surface of  $\text{TiO}_2$  particles is substantially less than in the volume. This is conditioned by high diffusion coefficient of Co atoms from the surface into  $\text{TiO}_2$  volume, which is observed even at low temperatures [6].

### 3.2 Structural and elemental analysis of $\text{TiO}_2$ samples

Optical absorption spectra measured in the range of photon energies from 1,75 to 3,1 eV for the samples of pure and modified  $\text{TiO}_2$  rutile and anatase are shown in Fig. 1a, b in the form of the absorption coefficients  $\alpha$ .



**Fig. 1** – Absorption spectra of pure and modified  $\text{TiO}_2$  in the range of 1,75-3,1 eV: 1 – R, 2 – R/Cr, 3 – R/Cu, 4 – R/Co, 5 – R/Fe (a) and 1 – A, 2 – A/Cu, 3 – A/Co, 4 – A/Fe, 5 – A/Cr (b)

For pure  $\text{TiO}_2$  in the energy region from 2,0 to 3,0 eV one can observe a weak absorption induced by optical transitions between the states positioned above the top of the valence band (VB) and lower the bottom of the conduction band (CB). As known, these states are connected with bulk and surface defects and can be populated due to thermal excitation of CB electrons even at room temperature [7]. Increase in the absorption observed for the samples of modified  $\text{TiO}_2$  can be connected with both the appearance of additional impurity states in the band gap and the intrinsic absorption of impurity atoms. These additional states are located above the top of the VB on 0,35 eV for Cr [7, 8] and 0,2 eV for Fe [2], on the top of the VB for Co [2], and within the BG for Cu [8], and all transitions lead to the sp-d charge transfer from these states to the CB of  $\text{Ti}^{4+}$ .

Obtained TiO<sub>2</sub> absorption spectra have the following peculiarities.

1. For all samples of modified anatase A/Co, A/Fe, and A/Cr (excluded A/Cu), substantial red shift of the absorption edge in the range 2,7-3,1 eV in comparison with pure anatase is observed. This shift correlates with the relative position of impurity levels in the TiO<sub>2</sub> band gap. The most expressed red shift is observed for A/Cr. Absorption peak near 2,80-2,85 eV (Fig. 1b) is connected with the increase in structural defects of anatase in comparison with rutile.

2. For all samples of rutile, in contrast to anatase, shift of the absorption edge is not observed. Absence of this shift implies that adsorption of impurities on rutile does not lead to the appearance of additional states in the TiO<sub>2</sub> band gap.

3. Absorption coefficients for rutile samples decrease as  $R < R/Cr < R/Cu < R/Co \leq R/Fe$ , and for anatase ones as  $A < A/Cu < A/Co < A/Fe < A/Cr$ .

4. Absorption curves for R/Fe and R/Co have almost the same behavior (Fig. 1a and Fig. 1b). This is connected with the fact that concentrations of Fe<sup>2+</sup> and Co<sup>2+</sup> impurity ions in these samples are almost equal, cations have identical ion radiuses and close intrinsic absorption bands (1,9-3,1 eV [5, 9] and 2,25-3,1 eV [10], respectively).

5. As for a small absorption for R/Cr, this is connected with low concentration of Cr<sup>3+</sup> ions in rutile in comparison with other impurities. Large absorption for A/Cr is connected with the increase in the concentration of Cr<sup>3+</sup> ions in anatase (Table 1) and strong intrinsic absorption of Cr at 2,0 eV [11-13] and 2,65 eV [14]. This corresponds to the transitions with charge transfer  $Cr^{3+} \rightarrow Ti^{4+}$  and  $4A_{2g} \rightarrow 4T_{1g}$  in octahedral environment [15].

6. Though concentration of Cu<sup>2+</sup> ions in A/Cu and R/Cu samples was the largest, absorption of these samples in the range 1,8-3,0 eV is the lowest. This is explained by the fact that intrinsic absorption of Cu is shifted to the region 2,81-3,87 eV [16] in comparison with other impurities.

Absorption spectra measured in the range of photon energies from 2,7 to 5,5 eV for samples of pure and modified TiO<sub>2</sub> rutile and anatase are shown in Fig. 2a, b. As seen, the influence of impurities on absorption is stronger expressed for anatase than for rutile, but the main laws are similar for all samples.

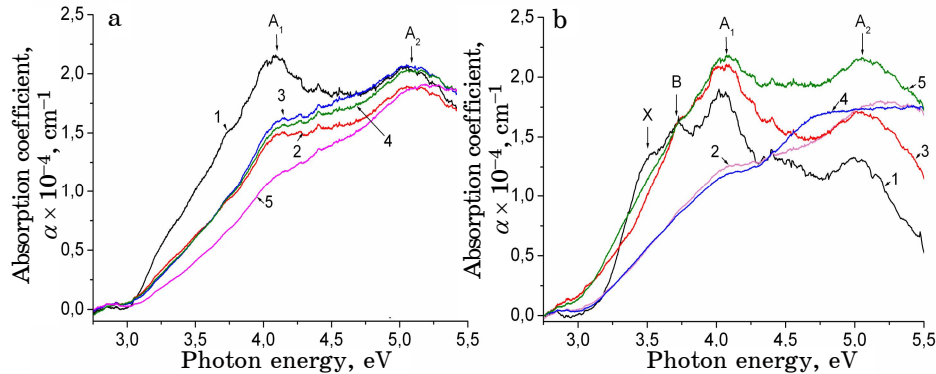


Fig. 2 – Absorption spectra of pure and modified TiO<sub>2</sub> in the range 2,7-5,5 eV: 1 – R, 2 – R/Cr, 3 – R/Cu, 4 – R/Fe, 5 – R/Co (a) and 1 – A, 2 – A/Co, 3 – A/Cr, 4 – A/Cu, 5 – A/Fe (b)

Two absorption bands with maximums  $A_1$  and  $A_2$  are observed in both pure and modified rutile and anatase. These bands correlate well with the bands in absorption spectra of monocrystalline  $TiO_2$  obtained experimentally [17] and calculated theoretically [18].

In absorption spectra of pure rutile (Fig. 2a) bands  $A_1$  and  $A_2$  are connected with the splitting of O  $2p_{x,y}$  orbital of oxygen atoms in the VB [19, 20]. Band  $A_1$  at 4,05 eV is connected with the transitions between upper edge of the VB and bottom of the CB (conditionally “first order” transitions). Band  $A_2$  at 5,08 eV is related to the transitions between four upper valent states of the VB and first six states  $t_{2g}$  in the CB (“second order” transitions). In absorption spectra of pure anatase (see Fig. 2b) bands  $A_1$  and  $A_2$  are located close to 4,03 and 5,00 eV, respectively, and are the consequence of the transitions from the VB into the CB. Also, absorption bands X and B at 3,5 and 3,7 eV, respectively, are observed for pure anatase. Band B was observed earlier for anatase monocrystal at 3,6 eV [21], but we did not find any information elsewhere about band X. Asahi et al. [22] ascribed band B to  $O^{2-} \rightarrow Ti^{4+}$  transitions with charge transfer. They have also shown that this transition is the dipole-forbidden at  $E \parallel c$  polarization and is the dipole-allowed at  $E \perp c$ .

We should notice the following peculiarities of  $TiO_2$  absorption spectra.

a) Impurities in rutile lead to the significant decrease in the absorption band  $A_1$  and the increase in the band  $A_2$ , and, as a consequence, to the redistribution of the absorption ratio in bands  $A_1$  and  $A_2$ . Minimum absorption of  $A_1$  is observed for R/Co sample. Absorption ratio in bands  $A_1$  and  $A_2$  for rutile R : R/Cr : R/Cu : R/Co : R/Fe is changed as 1 : 0,8 : 0,8 : 0,8 : 0,7. This implies the increase in the relative intensity of the “second order” transitions, i.e. between upper valent states of the VB and  $t_{2g}$  states in the CB.

b) Impurities in anatase lead to complex changes in the relative absorption in bands  $A_1$  and  $A_2$ . Absorption of  $A_1$  for doped anatase samples in comparison with pure anatase increases for A/Cr and A/Fe, but decreases substantially for A/Co and A/Cu. Absorption ratio A : A/Cr : A/Fe : A/Co : A/Cu in bands  $A_1$  and  $A_2$  is changed as 1,4 : 1,2 : 1,0 : 0,7 : 0,7. This implies the different effect of impurities on the electron transitions in anatase: relative intensity of the “first order” transitions increases for A/Cr and A/Fe and essentially decreases for A/Co and A/Cu. But for all samples of modified anatase relative intensity of the “second order” transitions substantially increases in comparison with pure anatase.

c) A wide background in the range 4,4-4,9 eV and shift of the band  $A_2$  to the long-wave region in comparison with other anatase samples are observed on the absorption spectra of A/Cu. This occurs due to the overlap of the absorption bands, which correspond to  $O^{2-}(2p) \rightarrow Ti^{4+}(3d)$  and  $O^{2-}(2p) \rightarrow Cu^{2+}(3d)$  transitions with charge transfer [23].

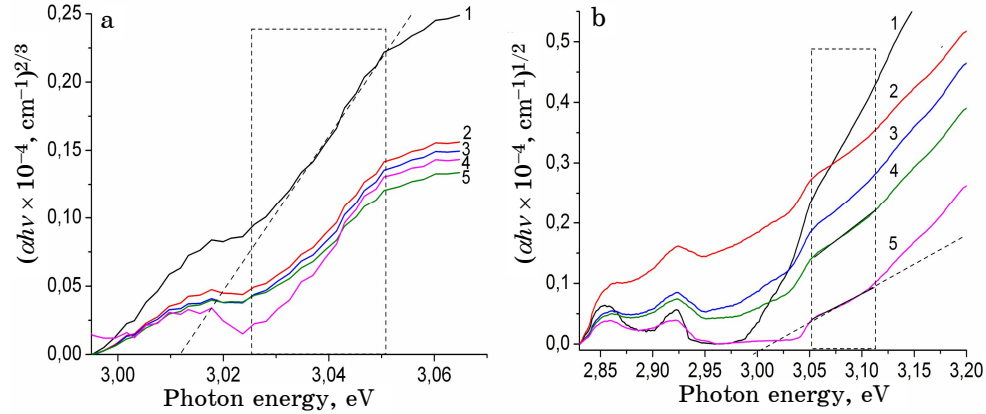
### 3.3 Band gap width

Band gap width  $E_g$  of semiconductors is connected with the absorption coefficient  $\alpha$  and can be determined from equation [24]

$$A (h\nu - E_g)^r = \alpha h\nu, \quad (1)$$

where  $h\nu$  is the photon energy;  $A$  is the constant, which correlates with the ordering of the crystal structure;  $r = 2$  for indirect allowed transitions;  $r = 3/2$

for direct forbidden transitions, and  $r = 1/2$  for direct allowed transitions. Band gap width of TiO<sub>2</sub> rutile and anatase samples was defined by the intersection of tangent  $(\alpha h\nu)^{1/r}$  to the photon energy axis as it is shown in Fig. 3 and Fig. 4, respectively.



**Fig. 3** – Dependence of  $(\alpha h\nu)^{1/r}$  on the photon energy for pure and modified rutile:: 1 – R, 2 – R/Cr, 3 – R/Cu, 4 – R/Co, 5 – R/Fe (a) and 1 – R, 2 – R/Cr, 3 – R/Cu, 4 – R/Fe, 5 – R/Cr (b)

Obtained values of the band gap width of direct  $E_{gd}$  and indirect  $E_{gi}$  transitions for polycrystalline rutile samples in comparison with the references are represented in Table 2.

**Table 2** – Values of the band gap width for rutile samples

Samples	$E_{gd}$ , eV $r = 1,5$	$E_{gi}$ , eV $r = 2$	$E_g$ , eV taken from the references (impurity content, %)	References
R	3,01	2,97	3,0	[29]
R/Fe	3,02	2,83	2,8 (1 at.%)	[29]
R/Cr	3,02	2,91	3,25 (1 at.); 3,15(10 at.)*	[30]
R/Co	3,03	2,99	2,72 (3 at.)*	[31]
R/Cu	3,02	2,93	3,00 (0,5 M.); 3,06 (1,0 M.)	[32]

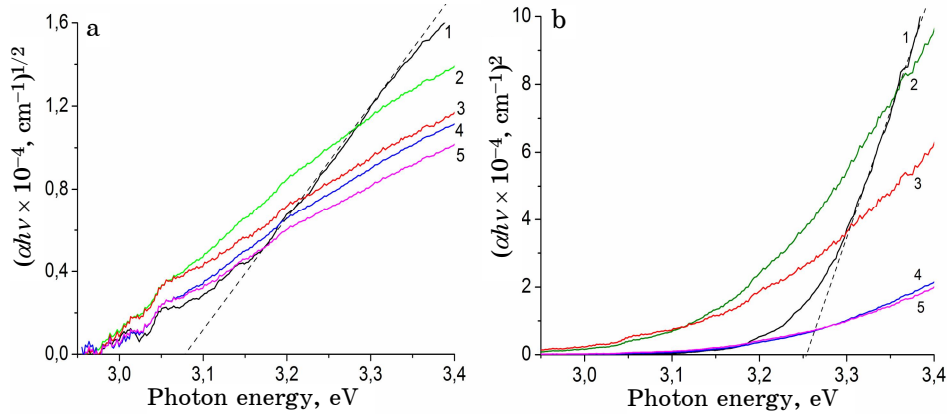
\* data for TiO<sub>2</sub> films.

For the samples of pure rutile we have found the band gap width of direct transitions:  $E_{gd} = 3,01$  eV. This value agrees well with the data for rutile monocrystal [27] and it was discussed in our previous work [28]. As follows from the experimentally obtained data, for modified rutile samples, in comparison with pure rutile, band gap width remains practically unchanged.

Energy states in the band structure of anatase, in contrast to rutile, have two maximums in the upper region of the VB, and difference in the energies between these two maximums is very small (0,1 eV) [22, 33]. Electron transition from these maximums to the bottom of the CB can be both direct  $\Gamma$  (VB)  $\rightarrow$   $\Gamma$  (CB) [22] and indirect M (VB)  $\rightarrow$   $\Gamma$  (CB) [33]. What this transition is – direct or indirect – this completely depends on the crystal structure, lat-

tice parameters, and dispersed composition of semiconductor. Though monocrystalline  $\text{TiO}_2$  is an indirect semiconductor,  $\text{TiO}_2$ -based nanostructured materials can exhibit properties of “direct” semiconductors [34, 35].

Starting from our experimental data, for polydisperse anatase samples we have determined the band gap width for two types of transitions: direct and indirect. In Fig. 4 we show the dependences  $(\alpha h\nu)^{1/r}$  on  $h\nu$  for anatase, which correspond to the direct and indirect transitions at  $r = 2$  (a) and  $r = 0,5$  (b), respectively.



**Fig. 4** – Dependence of  $(\alpha h\nu)^{1/r}$  on the photon energy for pure and modified anatase: 1 – A, 2 – A/Fe, 3 – A/Cr, 4 – A/Cu, 5 – A/Co (a) and 1 – A, 2 – A/Fe, 3 – A/Cr, 4 – A/Cu, 5 – A/Co (b)

Obtained values of the band gap width of direct  $E_{gd}$  and indirect  $E_{gi}$  transitions for polycrystalline anatase samples in comparison with the references are represented in Table 3.

**Table 3** – Values of the band gap width for anatase samples

Samples	$E_{gd}$ , eV $r = 0,5$	$E_{gi}$ , eV $r = 2$	$E_g$ , eV taken from the references (impurity content, %)	References
A	3,25	3,08	3,22; 3,19; 3,45; 3,29*	[36][37][32][39]
A/Cr	3,14	2,94	–	–
A/Cu	3,21	2,97	3,15 (0,5 M.%); 3,37 (1,0 M.%)	[32]
A/Co	3,20	2,94	3,26-2,53 (0-35%)* 3,14 (4 at.%), 2,96*(7 at.%)	[38] [37]
A/Fe	3,13	2,96	2,88 (4,89 at.) 2,83 (25 at.)* 3,15 (2 at.%)	[9] [39] [37]

\* data for  $\text{TiO}_2$  films.

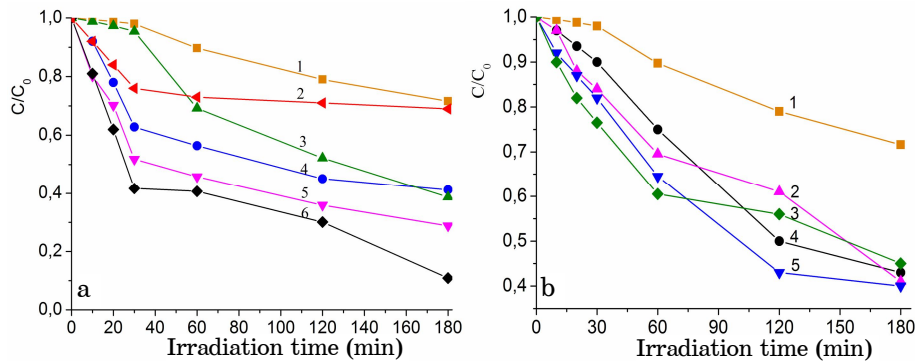
We have found some papers connected with indirect band gap for polydisperse  $\text{TiO}_2$ , and only the authors of [9] represented data for direct band gap of anatase. In the cases when we have not found data for  $\text{TiO}_2$  powders, data for  $\text{TiO}_2$  films was taken.



Obtained value of the band gap width for direct transitions in pure anatase  $E_{gd} = 3,25$  eV agrees well with the data of [36, 37] for  $\text{TiO}_2$  with the same particle size as in our samples. Our estimation of the direct band gap for A/Cu and A/Fe also correlates well with the data of [32, 37] for anatase with the same particle size and impurity concentration.

### 3.4 Photocatalytic activity

Reaction of photodestruction of organic dye safranin under the action of UV irradiation ( $\lambda_{\text{max}} = 254$  nm) was studied to estimate the photocatalytic activity of pure and modified anatase and rutile. In Fig. 5 we show the dependence of the relative safranin concentration in solution on the UV irradiation time.



**Fig. 5** – Photodestruction of safranin: (a) 1 – without catalyst, 2 – R/Cr, 3 – R/Fe, 4 – R/Cu, 5 – R/Co, 6 – R; (b) 1 – without catalyst, 2 – A/Co, 3 – A/Fe, 4 – A, 5 – A/Cu

In the absence of catalyst, photodecomposition of safranin is very slow, but the process is substantially accelerated in the presence of anatase or rutile, especially during first 30-60 minutes. The largest rate of safranin photodestruction was observed for pure rutile: in 30 minutes of UV irradiation almost 60% and in 180 min – about 90% of the dye are decomposed. Inhibition of photocatalytic reaction was observed for modified rutile. Modification by the transition metal cations of anatase, in contrast to rutile, leads to the increase in the reaction rate of safranin photodestruction. Thus, after 60 min of UV irradiation, relative concentration of safranin decreased: 25% for pure anatase, 30% for A/Co, 35% for A/Cu, 40% for A/Fe. Almost the same degree of dye photodecomposition (55-60%) was observed after 180 min of UV irradiation for all anatase samples.

Obtained results of the safranin photodestruction in the presence of  $\text{TiO}_2$  correlate well with the data concerning the photoabsorption of anatase and rutile. As known, the rate of photocatalytic reaction depends on many factors: crystal size, specific surface of particles, impurity concentration, etc. Relation between the rate of photoinduced charge transfer from the volume to the surface of  $\text{TiO}_2$  particle and the recombination rate of photoexcited electrons and holes plays an important role. Transition metal ion impurities, depending on their concentration, can be additional traps of photogenerated electrons and holes or be the centers of their recombination. There is an optimal concentration of the transition element impurities, above which these impurities act as the recombination centers [3].

Our experimental data shows that impurities in rutile (Fig. 5a) lead to the inhibition of safranin photodecomposition, because a large amount of these impurities leads to the appearance of additional recombination centers of the photoexcited electron-hole pairs. For anatase samples (Fig. 5b), inhibition of safranin photodecomposition due to the additional recombination centers is compensated by the increase in the rate of photocatalytic reaction because of the increase in the photoabsorption of modified anatase. Therefore, one can observe the correlation between the relative rate of photocatalysis and the relative coefficient of the optical absorption of anatase samples in the range 4,0-5,5 eV.

#### 4. CONCLUSIONS

The influence of Cu, Fe, Co, and Cr transition metal ion impurities on the optical and photocatalytic properties of polydisperse powders of nanocrystalline TiO<sub>2</sub> of rutile and anatase modifications with a high degree of chemical purity synthesized in the same experimental conditions is studied in the present work.

Optical absorption spectra of pure and modified samples of anatase and rutile in a wide range of photon energies (1,75-5,5 eV) are investigated.

It is shown that in the region 2-3 eV absorption of TiO<sub>2</sub> is connected with optical transitions between defect states localized between the top of the VB and the bottom of the CB. Substantial increase in the absorption for all TiO<sub>2</sub> samples is conditioned by absorption of impurity ions.

Spectral shift of the absorption edge was not observed for all rutile samples, while significant red shift of the absorption edge, which correlates well with the position of impurity levels in the TiO<sub>2</sub> band gap, was registered for doped anatase samples (except of A/Cu).

Absorption of TiO<sub>2</sub> in the region 3,0-5,5 eV is conditioned by the electron transitions between the VB and t<sub>2g</sub> states of the CB. Influence of impurities in rutile is observed only for the transitions in the region 3,0-4,5 eV, while absorption of doped anatase is substantially changed in the region 3,0-5,5 eV in comparison with pure anatase.

The band gap width of polydisperse TiO<sub>2</sub> samples is determined. It remains practically unchanged for doped rutile, whereas for doped anatase band gap width decreases being the largest for A/Cr and A/Fe samples.

Photocatalytic activity of TiO<sub>2</sub> samples in the reaction of photocatalytic decomposition of organic dye safranin under UV irradiation is studied. It is shown that for modified rutile the rate of photocatalytic reaction slows down because at large concentration of impurities they become the recombination centers of the photoexcited electron-hole pairs. On the contrary, for modified anatase acceleration of photocatalytic reaction is observed. Relative rate of photocatalysis in the presence of modified anatase correlates well with the relative coefficient of optical absorption of anatase samples in the region of 4,0-5,5 eV.

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